

Properties of magnesium oxide boards used as sheathing in exterior walls

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Abstract. Sheathing boards based on magnesium oxide (MgO) became very popular in the Danish market around the year 2010. However, it was realized a few years later that they condensed humidity from ambient air. The leakage from the boards led to severe problems with corrosion of metallic components and moisture uptake and degradation of wooden members. The boards have been banned from use in the Danish market since 2015, while damage cases valuing around 370 million EUR still need to be reconciled [1]. The binder in MgO board is formed by a chemical reaction between MgO and MgCl₂. However, MgSO₄ is a possible alternative to MgCl₂, and such boards may have smaller moisture uptake, to reduce this issue. This paper gives an account of recent activities to compare properties of MgO boards, based on chlorides or sulphates. The investigations comprise: (1) Analysis of elemental composition and crystalline components, (2) Examination in optical microscope and scanning electron microscope, (3) Determination of moisture retention and water vapour permeability, (4) Analysis of chemical composition of leaked salt water from boards, and (5) Visual observation of decomposition and determination of dry mass change after exposure to high humidity.

1 Introduction

Around the year 2010, a new sheathing board product, the magnesium oxide (MgO) board, became a popular building material in Denmark to be used inside ventilated façades as sheathing behind a ventilated air gap. However, during winter 2015 Ingeniøren – the official magazine of the Engineering Association of Denmark – began reporting widespread moisture damage inside ventilated facades of Danish buildings constructed with MgO-boards [2].

The binder in MgO-boards is formed by a chemical reaction between MgO and MgCl₂, which was first produced by Stanislas Sorel in 1867 and is known as Sorel cement or magnesium oxychloride cement. The product is hard and strong but decomposes over time by contact with water or air at high relative humidity [3].

An investigation of Danish buildings revealed that the MgO-boards suffer from the same durability issue as its binder. When MgO-boards are exposed to high humidity, they

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absorb an excessive amount of moisture from the air and form water drops that contain a high amount of chloride ions. This process is also referred to as “crying” [4].

To improve the resistance to degradation of the product, MgO-boards containing MgSO₄ have been introduced to the market because of the less hygroscopic behaviour of magnesium sulphate compared to magnesium chloride [5]. The purpose of this paper is therefore to investigate the moisture properties of the different types of MgO-boards.

2 Materials for investigation

All investigations were performed on the same three MgO-boards delivered by a Chinese manufacturer, where #1 (9 mm) and #2 (12 mm) are magnesium sulphate boards and #3 (12 mm) is a magnesium chloride based board. The elemental composition of the MgO-boards has been determined using both X-ray Fluorescence (XRF, Ameteks Spectro xSort x-Ray Spectrometer) and Energy Dispersive X-ray (EDX, Jeol JSM-6510 with Bruker XFlash Detector 410-M). Both analyses were performed on a cross-section of the materials. The results are presented in Table 1 where the main difference is the chloride and sulphate content.

Table 1. XRF and EDX results of the MgO-boards. Note that oxygen and carbon cannot be included in the measurement. The rest has been scaled to % of mass and represent an approximation of the binder composition.

	XRF results					EDX results				
	Mg	Si	S	Cl	Ca	Mg	Si	S	Cl	Ca
#1 – MgSO ₄ 9 mm	73%	10%	13%	2%	2%	84%	5%	9%	0%	1%
#2 – MgSO ₄ 12 mm	79%	6%	12%	2%	2%	80%	8%	10%	0%	2%
#3 – MgCl ₂ 12 mm	68%	9%	1%	21%	2%	77%	2%	0%	19%	1%

The crystalline components in the MgO-boards have been identified qualitatively using powder X-ray Diffraction (XRD) on samples with grain sizes less than 200µm on the X’Pert PRO PANalytical device with measurements in steps of 0.001° from 3° to 65°. The peaks were identified based on the “ICDD PDF-2/Release 2009 RDB” database and the results are shown in Table 2. Several hydrate phases form in these binders, which are in agreement with Walling & Provis [5]. SiO₂ and MgCO₃ are not hydration products and thus likely to originate from filler materials.

Table 2. Observed phases in the three MgO-boards using powder XRD.

	Chemical Formula	Ref. Code (PDF)	Comments
#1 – MgSO ₄ 9 mm	Mg(OH) ₂	00-001-1169	
	Mg ₄ (OH) ₆ SO ₄ · 8 H ₂ O	00-007-0418	Binders
	Mg ₆ (OH) ₁₀ SO ₄ · 3 H ₂ O	00-007-0415	
	MgCO ₃ ^{**}	01-080-0042	
	SiO ₂ ^{**}	01-076-0940	
#2 – MgSO ₄ 12 mm	Mg(OH) ₂	00-001-1169	
	Mg ₄ (OH) ₆ SO ₄ · 8 H ₂ O	00-007-0418	Binders
	MgCO ₃ ^{**}	01-080-0042	
	SiO ₂ ^{**}	01-076-0940	
#3 – MgCl ₂ 12 mm	Mg ₃ (OH) ₅ Cl · 4H ₂ O	00-007-0420	Binder
	Mg ₆ Al ₂ (CO ₃)(OH) ₁₆ · 4H ₂ O	00-054-1029	

^{**} Small quantities.

3 Microstructure analysis of MgO-boards

For the microstructural analysis of the boards, an optical microscope with plane- and crossed polarized light was used. The analysis revealed a high content of air in both types of MgO-boards. Furthermore, many different filler materials are present, of which some are highly porous, and perlite is most likely one. See Figure 1.

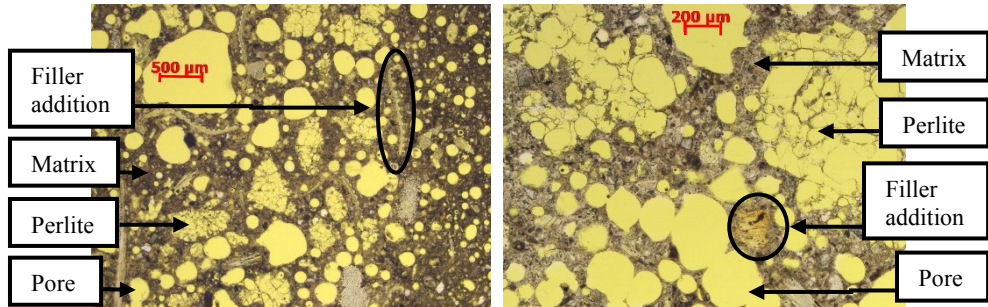


Fig. 1. Left picture: Specimen #1 - MgSO₄ 9 mm. Right picture: Specimen #3 - MgCl₂ 12 mm.

The specimens were further analysed using Scanning Electron Microscopy (SEM) with EDX in backscattered electron mode (BSE) with an FEI Quanta 200 ESEM microscope and an acceleration voltage of 15 keV. The EDX maps were measured with an Oxford Instrument (OI) X-Max^N detector and show that the matrix of the sulphate based boards consists of magnesium and sulphate with no visible content of chloride i.e. no contrast. Furthermore, the analysis revealed the matrix of the chloride based board to consist mostly of magnesium and chloride with no visible contrast for the sulphate. See Figure 2.

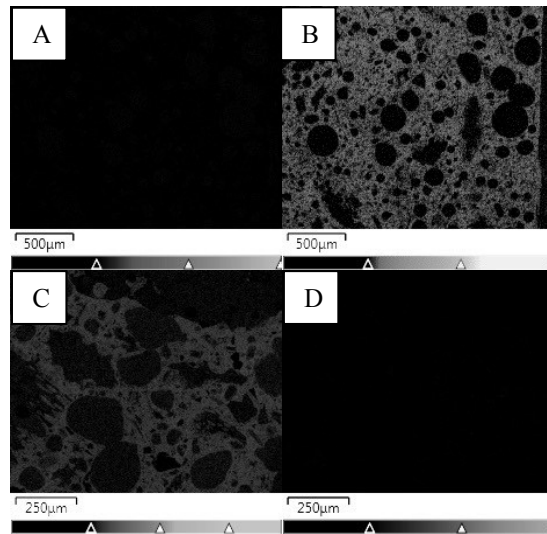


Fig. 2. Energy-Dispersive X-ray spectroscopy (EDX) maps of (A) chloride and (B) sulphate in specimen #1 - MgSO₄ 9 mm and of (C) chloride and (D) sulphate in specimen #3 - MgCl₂ 12 mm.

3.1 Qualitative evaluation of the volume percent composition

A quantitative evaluation of the volume percent composition was performed by counting a total of 1000 points per specimen at 50 times magnification through an ocular grid plate.

The data derived showed significant changes in the volume percentage up until 500 points, indicating the samples to be highly inhomogeneous. The results are presented in Table 3.

Table 3. Volume percentage of (1) air, (2) organic filler additions and perlite, and (3) the matrix.

	Air	Organic + perlite	Matrix
#1 – MgSO ₄ 9 mm	15%	15%	70%
#2 – MgSO ₄ 12 mm	20%	15%	65%
#3 – MgCl ₂ 12 mm	15%	20%	65%

4 Moisture retention measurements on MgO-boards from air

A moisture retention test was performed for seven weeks at 20 °C by placing samples of each MgO-board in a climate chamber. Before starting the experiment, a dry mass of about 10 g was obtained by placing the samples in a heating chamber at 105 °C for four days and cooled to room temperature in a desiccator. The samples were placed in petri glass dishes together with two watch glasses with the purpose of allowing ventilation and making it easy to observe any water drops forming underneath. The samples were then exposed to a relative humidity from 85% to 95% in steps of 2% RH.

Weightings were performed without opening the climate chamber by using a glove that is attached to the front door. Results from every step during the seven weeks are depicted in Figure 3 where the increase in mass is measured after seven days at all steps except 91% RH, which was kept for 14 days. Note that the boards are not in equilibrium with the humidity after seven days and would absorb more water given more time. Initial “tears” were observed on the chloride based board at 89% RH after having gained 43% in mass and mould growth was observed on it at 93% RH. The first sight of “tears” from the sulphate based boards was observed seven weeks into the experiment at 95% RH. At this point both sulphate based boards had gained around 60% in mass.

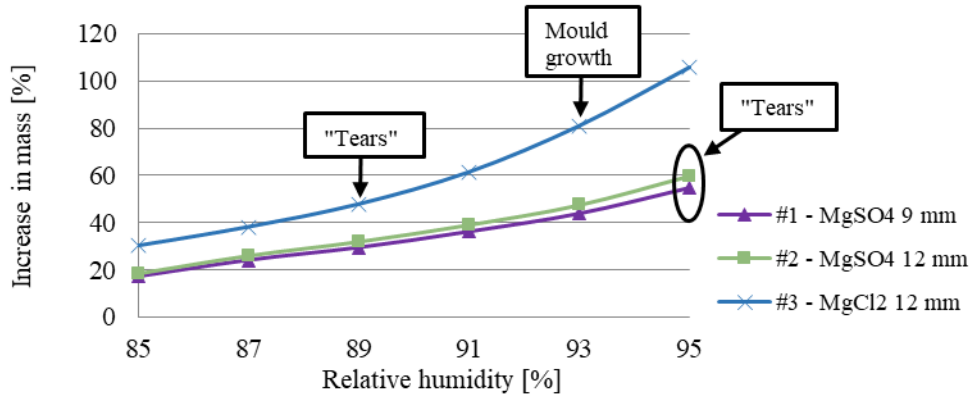


Fig. 3. Moisture retention curve from 85% RH to 95% RH measured over the course of seven weeks. The increase in mass includes the weight of “tears”.

The amount of “tears” and mould growth that had appeared on the chloride based board at the end of the moisture retention test can be seen in Figure 4 together with a picture of one of the sulphate based boards, which was kept at 95% RH for two additional weeks after the moisture retention test was over. The amount of “tears” and mould growth on the sulphate based boards was significantly less than that of the chloride based.

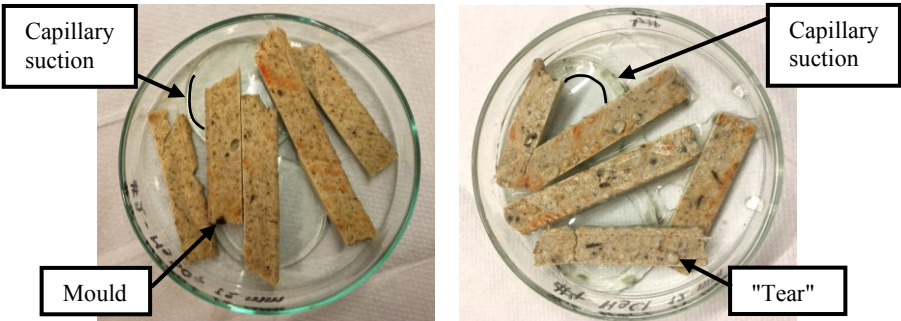


Fig. 4. Left picture: Sample #2 – MgSO₄ 12 mm after nine weeks of exposure to high humidity. Right picture: Sample #3 – MgCl₂ 12 mm after seven weeks of exposure to high humidity.

5 Water vapour resistance for MgO-boards

The water vapour resistance Z_p of the MgO-boards has been measured by using the ‘wet cup’ test described in EN 12572 with conditions of 50/93% RH and are presented in Table 4 along with the water vapour permeability δ_p [6]. The measured water vapour resistances are similar to the water vapour resistance of a 6 mm fibre cement board [7].

Table 4. Water vapour resistance Z_p and water vapour permeability δ_p .

	Water vapour resistance, Z_p [GPa·m ² ·s/kg]	Water vapour permeability, δ_p x 10 ⁻¹² [kg/(Pa·m·s)]
#1 – MgSO ₄ 9 mm	2.4±0.2	3.7±0.3
#2 – MgSO ₄ 12 mm	4.4±0.3	2.8±0.2
#3 – MgCl ₂ 12 mm	2.9±0.4	4.1±0.5

6 Chemical composition of leaked salt water from MgO-boards

At the end of the moisture retention test, “tears” were collected from the samples and an analysis of them was done using Inductively Coupled Plasma Optical Emission Spectrometry and Ion Chromatograph equipment. The results are presented in Table 5. For comparison, it can be informed that the Atlantic ocean contains 20000 Cl⁻ mg/liter [3].

Table 5. Results of ions found in ”tear” drops from MgO-boards given in mg/liter.

	Na ⁺	Mg ⁺⁺	S ⁻	Cl ⁻	K ⁺	Ca ⁺⁺
#1 – MgSO ₄ 9 mm	413	27794	36172	1490	269	323
#2 – MgSO ₄ 12 mm	406	28099	34721	1227	153	291
#3 – MgCl ₂ 12 mm	540	14053	753	44123	322	0

7 Visual observations and dry mass change after exposure to high humidity

The MgO-boards cannot be distinguished from each other by eye at time of delivery, as they are all the same shade of white. After one day at 85% RH, the chloride based board changed into a brownish colour, while the sulphate based boards turned brown after four weeks of exposure to a relative humidity between 85% and 91% RH.

After having collected “tears” from the samples, a new dry mass was measured by placing the samples in an oven at 105°C for four days and cooled to room temperature in a glass desiccator containing silica gel. When drying the boards, they once again turned white. The difference in dry mass for each sample is given in Table 6.

Table 6. Results of mass change relative to dry mass before moisture exposure.

	#1 – MgSO ₄ 9 mm	#2 – MgSO ₄ 12 mm	#3 – MgCl ₂ 12 mm
Mass change	+ 7%	+ 8%	- 0.7%

8 Discussion

The moisture retention measurements indicate the sulphate based boards to condense less water than the chloride based board tested when exposed to a high relative humidity. However, the MgSO₄-boards did start to “cry” at a relative humidity of 95% and the mass increased significantly at lower relative humidity – e.g. at outdoor RH levels which are common in a country such as Denmark. The obtained moisture properties of the sulphate based boards support references stating that magnesium oxysulphate cement suffers from susceptibility to degradation when exposed to moisture [5].

9 Conclusions

Based on the results, it can be concluded that both types of MgO-boards have an ability to absorb unacceptable amounts of water from a humid ambient, although the MgO-boards based on sulphate do condense less water from the environment than the MgO-board based on chloride when exposed to a high relative humidity. Thus, MgSO₄ is not a possible alternative that should be used in Denmark and countries with a similar humid climate.

References

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